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Reactant: The uranium(IV) oxide, UO_2 , which is used as fuel in nuclear power plants, has a higher percentage of the fissionable isotope uranium-235 than is present in the UO_2 found in nature. To make fuel-grade UO_2 , chemists first convert uranium oxides to uranium hexafluoride, UF_6 , whose concentration of uranium-235 can ...

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Exercise 10.2 - Limiting Reactant: The uranium(IV) oxide, UO_2 , which is used as fuel in nuclear power plants, has a higher percentage of the fissionable isotope uranium-235 than is present in the UO_2 found in nature.

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Chapter 10 - The Mole Section 10.1
Measuring Matter 1. pair 2. 5 3. dozen 4.
gross 5. 200 6. ream 7. 6,000,000,000 8.
0.5 mol 9. 6.02 $\times 10^{23}$ 10. four moles 11.
6.02 $\times 10^{23}$ Cu atoms 12. 4 $\times 10^{23}$ 13.
1 mol CH₄ 6.02 $\times 10^{23}$ molecules CH₄ 14.
23 2 ...

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Chapter 10 Study Guide. How many moles are in 34.6 L of BaO? How many grams are in 4.6 moles of PbSO₄? How many molecules are in 6.5 $\times 10^3$ moles of CO₂? If you have 2.68 moles of S₂O₃, how many atoms of S do you have? How many liters are in 4.39 $\times 10^{22}$ atoms of Mg? The compound methyl butanoate smells like apples.

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Flashcard maker : Lily Taylor. 1 test answers. Gram-atomic mass. the molar mass of an element expressed in grams per mole (g/mol) Gram-formula mass. The mass of a mole of formula units in a ionic compound bond. Mole. the simplest structural unit of an element or

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compound.

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Chapter 10: The Mole - ANNE SCHMIDT
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MAIN Idea Chemists use the mole to count atoms, molecules, ions, and formula units.

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